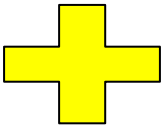
Directions: **Name:** _____

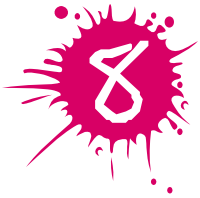
Place the definition next to the correct vocabulary word. Glue the definition into the correct box.

Ions	
Cation	
Anion	
Valence Shell	
Valence Electron	
Ionic Bond	

Covalent Bond	
Metals	
Non Metals	
Metalloids	
Octet Rule	
Oxidation Number	



Atoms that lose electrons and have a positive charge



Atoms want to have 8 electrons in their outer shell = stable

Atoms that gain electrons and have a negative charge

These are found to the left of the staircase and form ionic bonds with non-metals: give away electrons and become + Hydrogen is the exception.

Na^{+}

Cl^{-}

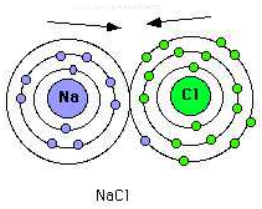
Atoms with either a positive (+) or negative (-) electrical charge

These are found to the right of the staircase and form covalent bonds with each other, or take electrons away from metals to form ionic bonds and become -

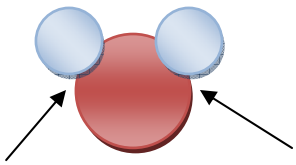
Ca^{+2}

The total number of electrons that an atom either gains or loses in order to form a chemical bond with another atom.

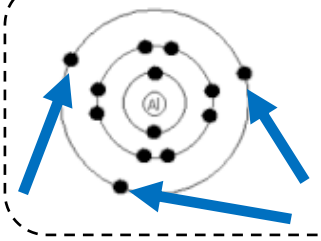
These touch the staircase in the periodic table and have properties of both metals and non-metals.



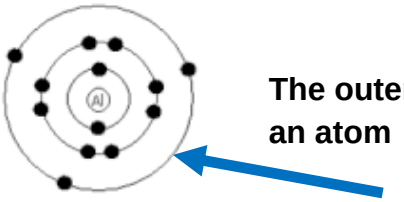
Type of bond where atoms give electrons to other atoms and there is an attraction between oppositely-charged ions



This is the type of bond where electrons are shared between atoms. H_20 – electrons are shared between H's & O.



The electrons found in the outermost shell of an atom. This is where the chemical reactions and bonding takes place.



The outermost shell of an atom